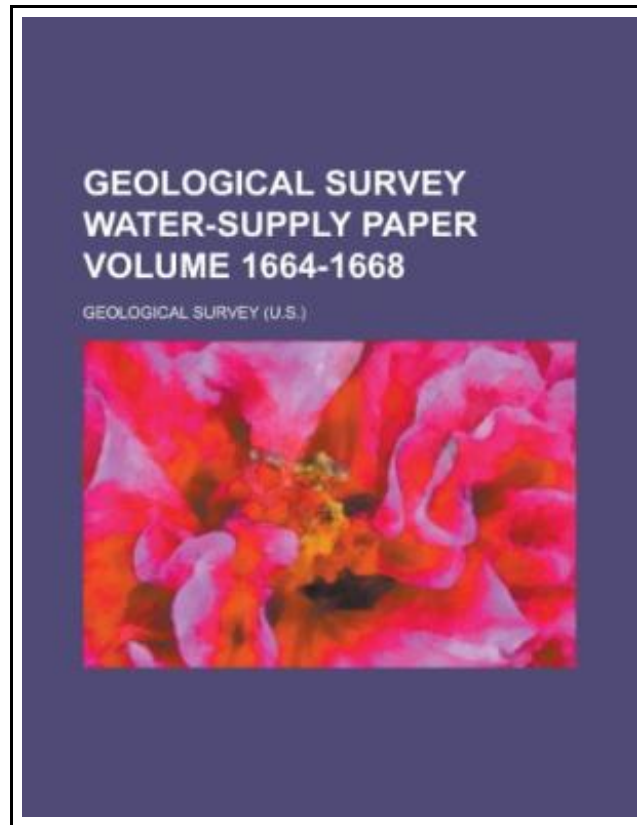


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Rarebooksclub.com, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. Excerpt: .solution were diluted to a manganous-ion concentration of 10 ppm. The portions were then raised to several selected pH values, ranging from 7.0 to 9.5, by adding sodium hydroxide. The solutions were then allowed to remain in contact with air, and samples were taken from time to time for determination of dissolved manganese by the standard periodate oxidation procedure. Before the manganese concentration was determined, the pH and potential of a bright platinum electrode, assumed to be the Eh of the sample, were measured, and the aliquot was filtered through a plastic-membrane filter whose pore diameter was 0.45 μ . Results are given in table 3. All these solutions tended to reach a constant pH near 7.0 after standing for a day or more. All the manganese was retained in solution except where the initial pH was 9.0 or greater. In most of the solutions the measured potential had a steady upward trend, and after 5 days it was between 0.54 and 0.56 v. The Mn^{+2} in these solutions was oxidized to Mn^{+3} or Mn^{+4} , in part, and these forms precipitated as Mn_2O_3 or MnO_2 , thereby lowering the pH. When the pH became unfavorable for further oxidation, the reactions stopped. In solution 2, measured and computed potentials remained close together for the entire series of observations (table 3). In solutions 3 to 6, a decided lack of equilibrium was observed at the start of the test. After 30 minutes, equilibrium had returned to solution 3, and...



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